

# Tổng hợp composite $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$ làm chất xúc tác quang hiệu năng cao phân hủy methylene blue

## TÓM TẮT

Trong nghiên cứu này, vật liệu composite  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  được tổng hợp bằng cách chuẩn bị riêng biệt hai thành phần  $\text{g-C}_3\text{N}_4$  và  $\text{Co}_2[\text{Fe}(\text{CN})_6]$ , sau đó kết hợp chúng thông qua quá trình phân tán  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  lên bề mặt  $\text{g-C}_3\text{N}_4$  ở  $90^\circ\text{C}$  trong 1 giờ. Các vật liệu thu được được đặc trưng bằng các kỹ thuật nhiễu xạ tia X (XRD), phổ hồng ngoại biến đổi Fourier (FT-IR), phổ tán xạ năng lượng tia X (EDS), hiển vi điện tử quét (SEM), phổ phản xạ khuếch tán UV-Vis (UV-Vis DRS) và phổ phát quang (PL). Hiệu suất xúc tác quang của vật liệu composite được đánh giá thông qua quá trình phân hủy methylene blue (MB) trong dung dịch nước dưới chiếu xạ ánh sáng mặt trời. Vật liệu composite  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  cho thấy hoạt tính xúc tác quang vượt trội so với từng thành phần riêng lẻ là  $\text{g-C}_3\text{N}_4$  và  $\text{Co}_2[\text{Fe}(\text{CN})_6]$ .

**Từ khóa:**  $\text{Co}_2[\text{Fe}(\text{CN})_6]$ ,  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$ , photocatalytic activity, methylene blue, sunlight

# Synthesis of $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$ composite as an improved photocatalyst for degradation of methylene blue

## ABSTRACT

In this study, the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite was synthesized by first preparing  $\text{g-C}_3\text{N}_4$  and  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  separately, followed by their coupling through the dispersion of  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  onto  $\text{g-C}_3\text{N}_4$  at  $90^\circ\text{C}$  for 1 hour. The resulting materials were characterized by X-ray diffraction (XRD), Fourier-transform infrared spectroscopy (FT-IR), energy-dispersive X-ray spectroscopy (EDS), scanning electron microscopy (SEM), UV-Vis diffuse reflectance spectroscopy (UV-Vis DRS), and photoluminescence (PL) spectroscopy. The photocatalytic performance of the composite was evaluated via the degradation of methylene blue (MB) in aqueous solution under sunlight irradiation. The  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite exhibited superior photocatalytic activity compared to the individual components,  $\text{g-C}_3\text{N}_4$  and  $\text{Co}_2[\text{Fe}(\text{CN})_6]$ .

**Keywords:**  $\text{Co}_2[\text{Fe}(\text{CN})_6]$ ,  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$ , photocatalytic activity, methylene blue, sunlight.

## 1. INTRODUCTION

Faced with the growing reality of seriously polluted water sources that threaten both public health and the environment, various technologies have been developed for wastewater treatment. However, conventional methods often fail to completely eliminate toxic contaminants and typically require substantial energy inputs.<sup>1</sup> To align with the principles of sustainable chemistry, it is essential to develop more efficient reaction conditions and advanced catalytic systems for pollutant removal, thereby ensuring water quality.<sup>2</sup> Photocatalysis is a light-driven chemical process that occurs when a photocatalyst is activated under light irradiation. This technique has attracted considerable attention due to its simplicity, cost-effectiveness, non-toxicity, high degradation efficiency, and long-term stability. It is widely recognized as a promising strategy for wastewater treatment and environmental protection.<sup>3</sup>

In recent years, semiconductor materials such as  $\text{TiO}_2$ ,  $\text{ZnO}$ ,  $\text{SnO}_2$ ,  $\text{g-C}_3\text{N}_4$ , and perovskites have been extensively investigated for photocatalytic applications in environmental remediation. Among emerging photocatalysts,  $\text{Co}_x[\text{Fe}(\text{CN})_6]_y$  has attracted growing attention and is gradually demonstrating its potential in photocatalysis, showing promise as a candidate for broad environmental applications.

Despite its outstanding advantages—such as strong visible-light absorption, non-toxicity to humans and other organisms, absence of secondary pollution, and ease of recyclability—

$\text{Co}_x[\text{Fe}(\text{CN})_6]_y$  exhibits a major limitation: its high solubility in aqueous media. This characteristic poses significant challenges for its synthesis and practical application.

In 2013, Dwivedi et al.<sup>4</sup> employed gel beads containing cobalt hexacyanoferrate immobilized on micrometer-sized alginate beads (a salt of alginic acid,  $(\text{C}_6\text{H}_8\text{O}_6)$ ) as an adsorbent material for the removal of radioactive contaminants—such as cesium, an element with a half-life spanning several decades—from aqueous solutions. In 2016, K. A. Lin et al.<sup>5</sup> successfully synthesized a series of Prussian blue analogues with the general formula  $\text{M}(\text{II})_3[\text{M}(\text{III})(\text{CN})_6]_2$  ( $\text{M}^{2+} = \text{Co}, \text{Cu}, \text{Fe}, \text{Mn}, \text{Ni}$ ;  $\text{M}^{3+} = \text{Co}, \text{Fe}$ ), which exhibited high photocatalytic activity. These materials were capable of efficiently degrading Rhodamine B (RhB) in aqueous solution when activated by peroxymonosulfate (PMS) under neutral conditions. Cobalt-based Prussian blue analogues (PBAs) have shown great potential as efficient catalysts for water-splitting reactions and the degradation of organic pollutants. Moreover, the Co–Fe PBAs retain their original crystalline structure even after multiple recycling cycles, indicating excellent structural stability.

Several studies have employed metal-containing complexes and their isomers as polymeric materials. These isomeric complexes often crystallize in cubic structures and feature metal ions in varying oxidation states.<sup>6</sup> The frameworks of these complexes possess porous structures, which endow them with a high specific surface area, excellent adsorption capacity, and numerous catalytic active sites. As a result, such

complexes have been extensively studied across various fields, particularly in photocatalysis.<sup>7</sup> Therefore, the modification of these complexes with semiconductors to develop novel materials exhibiting enhanced photocatalytic activity has attracted significant attention from researchers.

## 2. EXPERIMENTS

### 2.1. Chemicals

Chemicals were purchased from Sigma-Aldrich (Germany) with high purity, including cobalt sulfate ( $\text{CoSO}_4$ ,  $\geq 99.0\%$ ), potassium ferrocyanide ( $\text{K}_4[\text{Fe}(\text{CN})_6]$ ,  $\geq 98.5\%$ ), melamine ( $\text{C}_3\text{H}_6\text{N}_6$ ,  $\geq 99.0\%$ ), methylene blue hydrate ( $\text{C}_{16}\text{H}_{18}\text{ClN}_3\text{S}\cdot\text{xH}_2\text{O}$ ,  $\geq 97.0\%$ ), and ethanol ( $\text{C}_2\text{H}_5\text{OH}$ ,  $\geq 99.9\%$ ). All chemicals were used without further purification.

### 2.2. Materials preparation

**2.2.1. Synthesis of  $\text{g-C}_3\text{N}_4$ :** Five grams of melamine were ground in an agate mortar and then transferred to a porcelain crucible. The crucible was tightly covered with aluminum foil to minimize sublimation and enhance the condensation efficiency of  $\text{g-C}_3\text{N}_4$ . The sample was heated in a muffle furnace at  $550\text{ }^\circ\text{C}$  for 2 hours with a heating rate of  $10\text{ }^\circ\text{C}/\text{min}$ . After natural cooling to room temperature, the resulting product—yellowish in color—was collected and labeled as  $\text{g-C}_3\text{N}_4$ .

**2.2.2. Synthesis of  $\text{Co}_2[\text{Fe}(\text{CN})_6]$ :** 5 milliliters of  $0.2\text{ M}$   $\text{CoSO}_4$  solution were added to a  $200\text{ mL}$  glass beaker. Then,  $100\text{ mL}$  of  $0.05\text{ M}$   $\text{K}_4[\text{Fe}(\text{CN})_6]$  solution were gradually added at a rate of  $5\text{ mL}/\text{min}$  under continuous stirring at room temperature. A precipitate formed during the addition. The mixture was stirred for an additional 60 minutes, followed by ultrasonic treatment at  $90\text{ }^\circ\text{C}$  for 2 hours. The resulting solid was collected by centrifugation and washed repeatedly with distilled water until the pH of the filtrate reached 7. Finally, the sample was dried at  $80\text{ }^\circ\text{C}$  for 48 hours. The final product was identified as  $\text{Co}_2[\text{Fe}(\text{CN})_6]$ .

**2.2.3. Synthesis of  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite:** Cobalt sulfate ( $\text{CoSO}_4$ ) was dissolved in deionized water under stirring until fully dissolved. A suitable amount of  $\text{g-C}_3\text{N}_4$  powder was then added to the solution, followed by stirring and ultrasonication for 60 minutes at room temperature to promote the adsorption of  $\text{Co}^{2+}$  ions onto the  $\text{g-C}_3\text{N}_4$  surface. The resulting suspension was centrifuged to collect the solid. Subsequently, a  $\text{K}_4[\text{Fe}(\text{CN})_6]$  solution (with a mass ratio of  $\text{K}_4[\text{Fe}(\text{CN})_6]$  to  $\text{g-C}_3\text{N}_4$  of 1:3) was added to the solid under continuous stirring

at  $90\text{ }^\circ\text{C}$  for 60 minutes to allow in situ formation of the  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  framework on the  $\text{g-C}_3\text{N}_4$  matrix. The product was separated by centrifugation, washed thoroughly, and dried at  $80\text{ }^\circ\text{C}$  for 48 hours to yield a blue solid of  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$ .

### 2.3. Characterization

X-ray diffraction (XRD) patterns were recorded using a Bruker D8 Advance diffractometer equipped with a  $\text{Cu K}\alpha$  radiation source ( $\lambda = 1.540\text{ \AA}$ ), operating at  $30\text{ kV}$  and  $0.01\text{ A}$ . Fourier-transform infrared (FTIR) spectra were obtained using a PerkinElmer GX spectrometer. The morphology and elemental composition of the samples were examined by scanning electron microscopy (SEM) and energy-dispersive X-ray spectroscopy (EDS) using a Nova Nano SEM 450 system. UV–Vis diffuse reflectance spectra (UV–Vis DRS) were measured using a Jasco V-770 spectrophotometer, while photoluminescence (PL) spectra were recorded on a Horiba Jobin Yvon FL322 fluorescence spectrometer. The concentration of MB solution was determined by a photometric method using a Jenway 6800 UV–Vis spectrophotometer.

### 2.4. Photocatalytic activity

The adsorption equilibrium time of MB on the materials was investigated under dark conditions using the following procedure. A total of  $0.07\text{ g}$  of the material sample was added to a  $250\text{ mL}$  beaker containing  $120\text{ mL}$  of MB solution ( $10\text{ mg}/\text{L}$ ). The beaker was covered with a wooden enclosure to prevent light exposure and continuously stirred using a magnetic stirrer. At 30-minute intervals,  $8\text{ mL}$  of the suspension was withdrawn, centrifuged to separate the supernatant, and the resulting solution was transferred into a vial. The collected samples were labeled as  $t_0, t_1, t_2, \dots, t_8$ , corresponding to sampling times of 0, 30, 60, ..., and 240 minutes.

The adsorption capacity is calculated using the following equation:

$$q = \frac{(C_i - C_f) \cdot V}{m}$$

where  $q$ : adsorption capacity at a given time ( $\text{mg}/\text{g}$ );  $C_i$ : initial concentration of the solution ( $\text{mg}/\text{L}$ );  $C_f$ : concentration of the solution at the given time ( $\text{mg}/\text{L}$ );  $V$ : volume of the solution subjected to adsorption ( $\text{L}$ );  $m$ : mass of the adsorbent ( $\text{g}$ ).

A plot of adsorption capacity versus time was drawn to determine the time required to reach adsorption equilibrium. Following the adsorption

study, the sample was further exposed to sunlight to evaluate its photocatalytic activity. The degradation efficiency of MB by the materials was calculated using the following equation:

$$H = \frac{C_0 - C_t}{C_0} \times 100\%$$

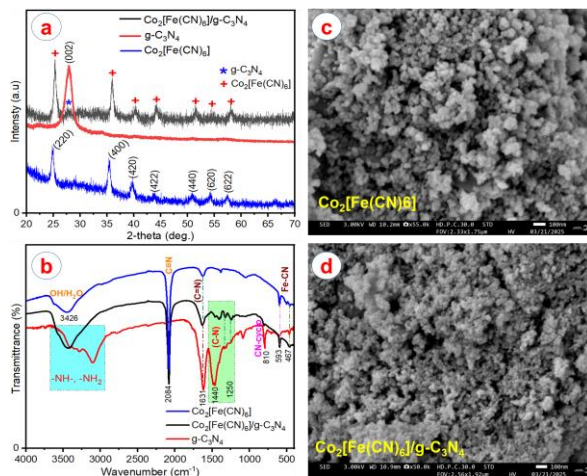
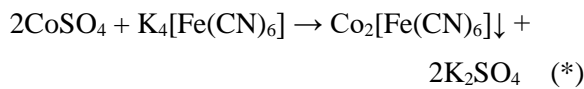
where  $C_0$  and  $C_t$  are the MB concentrations at the initial and selected  $t$ -times, respectively.

### 3. RESULTS AND DISCUSSIONS

#### 3.1. Material characterization

The  $\text{Co}_2[\text{Fe}(\text{CN})_6]$ ,  $\text{g-C}_3\text{N}_4$ , and  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  samples were characterized by XRD, FTIR, and SEM techniques.

As shown in Fig. 1a, the XRD pattern of the  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  sample displays distinct diffraction peaks at  $2\theta = 25.27^\circ$ ,  $35.92^\circ$ ,  $44.45^\circ$ ,  $51.53^\circ$ ,  $54.01^\circ$ , and  $58.38^\circ$ , which correspond to the (220), (400), (422), (440), (620), and (622) crystal planes, respectively. These results are in good agreement with previous reports on  $\text{Co}_2[\text{Fe}(\text{CN})_6]$ .  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  exhibits high crystallinity with characteristic diffraction peaks corresponding to the Prussian blue analogue (PBA) phase, in accordance with JCPDS card No. 77-1161.<sup>8-10</sup> The formation of the PBA-type  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  material from its precursors can be represented by the following reaction equation:



**Figure 1.** (a) XRD patterns; (b) FTIR spectra; (c, d) SEM images of  $\text{Co}_2[\text{Fe}(\text{CN})_6]$ ,  $\text{g-C}_3\text{N}_4$ , and the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite

The XRD pattern of pure  $\text{g-C}_3\text{N}_4$  exhibits a prominent diffraction peak at  $2\theta = 27.4^\circ$ , which corresponds to the (002) crystal plane. This peak is attributed to the interlayer stacking of conjugated aromatic systems in the graphitic carbon nitride structure. The  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-}$

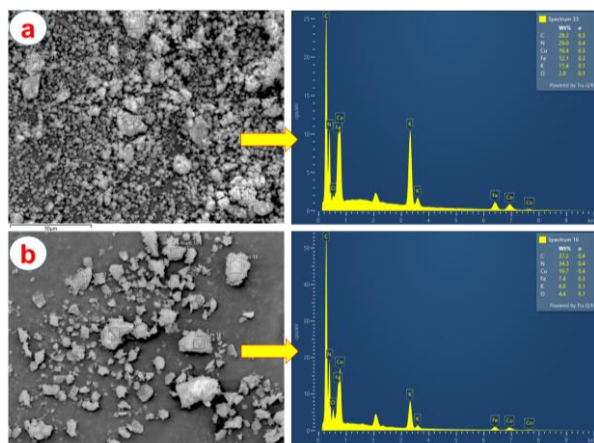
$\text{C}_3\text{N}_4$  composite clearly exhibits the characteristic diffraction peaks of both constituent materials,  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  and  $\text{g-C}_3\text{N}_4$ . Moreover, no additional crystalline phases are detected, indicating that the composite consists solely of these two components. This result is consistent with previously reported studies on PBA/ $\text{g-C}_3\text{N}_4$  materials.<sup>8</sup>

The FTIR spectra of the samples are presented in Fig. 1b. The  $\text{g-C}_3\text{N}_4$  sample exhibits a distinct absorption band at  $810 \text{ cm}^{-1}$ , which is attributed to the out-of-plane bending vibration of C–N bonds within the tri-s-triazine (heptazine) units—the fundamental structural motifs of the graphitic carbon nitride framework.<sup>11,12</sup> The sharp absorption bands in the region of  $1250\text{--}1631 \text{ cm}^{-1}$  are attributed to the stretching vibrations of C–N and C=N bonds within the aromatic rings of the heptazine-based structure.<sup>13,14</sup> In addition, a broad absorption band in the range of  $3100\text{--}3400 \text{ cm}^{-1}$  is observed, which is attributed to the stretching vibrations of –NH– and –NH<sub>2</sub> groups originating from the incompletely condensed structure of  $\text{g-C}_3\text{N}_4$ . The FTIR spectrum of  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  displays several characteristic absorption bands. A strong and sharp peak at  $2084 \text{ cm}^{-1}$  is attributed to the stretching vibration of the C≡N bond in the cyanide ligands, while a prominent band at  $593 \text{ cm}^{-1}$  corresponds to the bending vibration of the same group. Additionally, the absorption band at  $467 \text{ cm}^{-1}$  is ascribed to the stretching vibration of the Fe–CN bond,<sup>15,16</sup> confirming the metal–ligand coordination typical of Prussian Blue analogues (PBAs). Furthermore, a broad absorption band centered around  $3426 \text{ cm}^{-1}$  is associated with the –OH stretching vibration of water molecules adsorbed on the material surface. For the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite, the characteristic peaks corresponding to both  $\text{g-C}_3\text{N}_4$  and  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  were clearly observed, indicating the successful integration of the two components. In particular, the intense absorption band at  $2084 \text{ cm}^{-1}$  (assigned to the –C≡N stretching vibration) remained prominent, while the C–N and C=N stretching vibrations of  $\text{g-C}_3\text{N}_4$  were also retained. These observations confirm that the key structural features of both phases were preserved in the composite. Moreover, the absence of new absorption bands or significant peak shifts suggests that the interaction between  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  and  $\text{g-C}_3\text{N}_4$  is primarily physical in nature, possibly involving electrostatic attractions or hydrogen bonding. This result is in good agreement with the XRD data.



The SEM images of the  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  sample (Fig. 1c) and the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite (Fig. 1d) reveal distinct morphological characteristics. The pristine  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  exhibits a granular morphology composed of irregularly shaped, densely packed nanoparticles with a pronounced tendency toward agglomeration. The surface of the material appears relatively rough and compact, which may limit the accessibility of active sites for catalytic processes. In contrast, the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite exhibited a more porous architecture with significantly reduced particle agglomeration. The nanoparticles were more uniformly dispersed, with increased interparticle voids, likely attributed to the structural support provided by the  $\text{g-C}_3\text{N}_4$  matrix. This morphology suggests enhanced textural properties, such as a larger specific surface area and improved accessibility to active sites—factors that are critical for adsorption and photocatalytic performance. The morphological differences between the two samples provide clear evidence for the successful incorporation of  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  into the  $\text{g-C}_3\text{N}_4$  matrix. This integration effectively suppresses excessive particle agglomeration, enhances interfacial contact, and facilitates charge transfer during the photocatalytic process. These observations are consistent with the XRD and FTIR analyses, further confirming the formation of a structurally and functionally synergistic composite material.

The EDS spectra in Fig. 2 and Table 1 confirm the elemental composition of the samples.



**Figure 2.** EDS spectra of (a) the  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  sample and (b) the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite

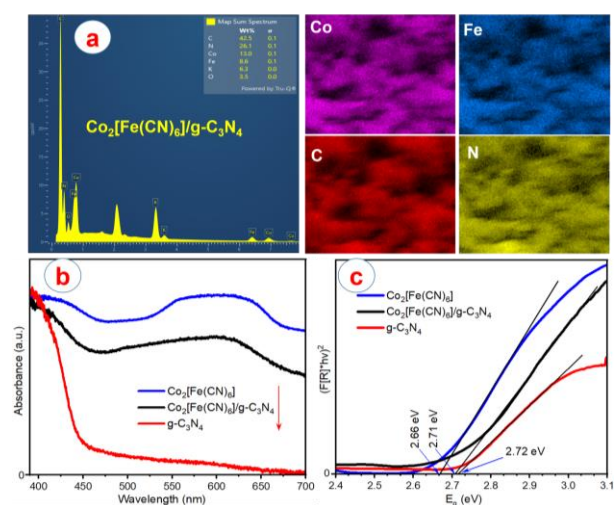
Both the  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  and  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  samples exhibit characteristic peaks of Co, Fe, C, and N, consistent with the expected structure of Prussian Blue analogues (PBA). Additionally, minor signals of K and O are observed. The presence of K is attributed to residual

$\text{K}_4[\text{Fe}(\text{CN})_6]$  precursor, while the O signal likely originates from surface-adsorbed moisture ( $\text{H}_2\text{O}$ ). In the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite, the increased intensities of C and N peaks are ascribed to the contribution of the  $\text{g-C}_3\text{N}_4$  matrix, while the decreased Co and Fe contents suggest effective dispersion of the PBA phase on the  $\text{g-C}_3\text{N}_4$  surface. These findings confirm the successful formation of the composite material.

**Table 1.** Elemental composition of  $\text{Fe}_3[\text{Fe}(\text{CN})_6]_2$  and  $\text{Fe}_3[\text{Fe}(\text{CN})_6]_2/\text{g-C}_3\text{N}_4$

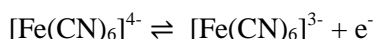
$\text{Co}_2[\text{Fe}(\text{CN})_6]$		$\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$	
Element	Weight%	Element	Weight%
C	29.2	C	37.2
N	29.0	N	34.3
Co	16.4	Co	10.7
Fe	12.1	Fe	7.4

The EDS-mapping (Fig. 3a) results show that Co, Fe, C, and N elements are uniformly distributed across the sample surface, indicating the successful integration and homogeneous dispersion of  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  within the  $\text{g-C}_3\text{N}_4$  matrix. No significant aggregation or phase separation was observed, suggesting strong interfacial compatibility between the two components. This uniform distribution supports the structural integrity of the composite and correlates well with the results obtained from XRD, FTIR, and SEM, reinforcing the successful synthesis of a well-dispersed hybrid photocatalyst.



**Figure 3.** (a) Elemental mapping images of Co, Fe, C, and N for the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite, (a) UV-Vis DRS spectra and (b) band gap energy of the materials

The UV–Vis DRS (Fig. 3b,c) reveal that all the materials exhibit absorption edges within the visible light region. Notably, the  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  and  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  samples show extended absorption edges reaching up to approximately 700 nm, which is consistent with the characteristic charge transfer transitions between Fe(II) and Fe(III):<sup>17</sup>



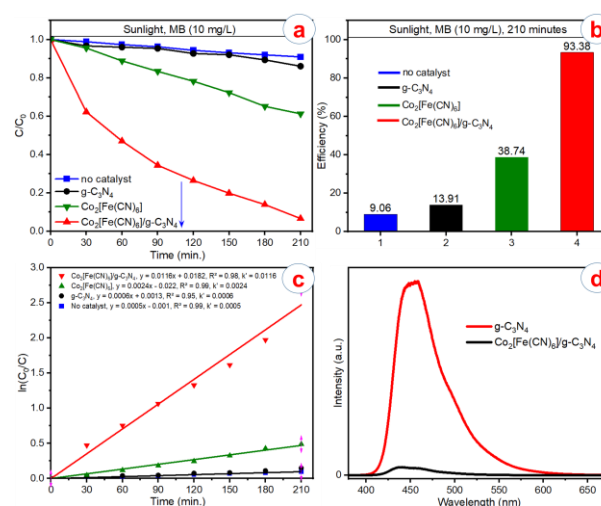
The band gap energy ( $E_g$ ) of the materials was also estimated using this method. Accordingly, the band gap values of  $\text{Co}_2[\text{Fe}(\text{CN})_6]$ ,  $\text{g-C}_3\text{N}_4$ , and the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite were determined to be 2.66 eV, 2.72 eV, and 2.71 eV, respectively.

### 3.2. Photocatalytic activity

The photocatalytic performance of the materials was evaluated through the degradation of MB under sunlight irradiation, with the results illustrated in Fig. 4a. After 210 minutes of exposure, the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite exhibited a degradation efficiency of 93.38%, outperforming both pristine  $\text{g-C}_3\text{N}_4$  and  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  (Fig. 4b). In addition, to compare the MB degradation rates among the materials, the photodegradation kinetics were analyzed (Fig. 4c and Table 2). Under identical illumination conditions, the degradation rate of MB using the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite was approximately five times higher than that of  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  and more than 19 times higher than that of pristine  $\text{g-C}_3\text{N}_4$ . This significant enhancement is attributed to the effective suppression of the rapid recombination of photogenerated electron–hole pairs, a common limitation in single-component  $\text{g-C}_3\text{N}_4$  photocatalysts, as confirmed by the photoluminescence (PL) spectrum shown in Fig. 4d.

**Table 2.** Results of MB degradation by the materials

Materials	Photodegradation (%)	Rate constant, $k$ ( $\text{min}^{-1}$ )	Regression coefficient ( $R^2$ )
$\text{g-C}_3\text{N}_4$	13.91	$0.06 \cdot 10^{-2}$	0.95
$\text{Co}_2[\text{Fe}(\text{CN})_6]$	38.74	$0.24 \cdot 10^{-2}$	0.99
$\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$	93.38	$1.16 \cdot 10^{-2}$	0.98



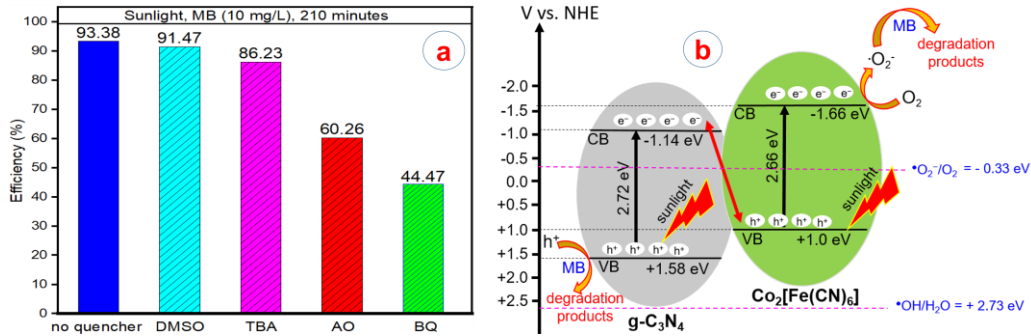
**Figure 4.** (a–c) Photocatalytic degradation of MB over  $\text{Co}_2[\text{Fe}(\text{CN})_6]$ ,  $\text{g-C}_3\text{N}_4$ , and  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$ ; (d) PL spectra of the corresponding materials

The photoluminescence (PL) spectra of  $\text{g-C}_3\text{N}_4$  and the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite were recorded under 365 nm excitation to evaluate the separation and recombination behavior of photoinduced electron–hole pairs ( $e^-/h^+$ ). In the spectrum, pristine  $\text{g-C}_3\text{N}_4$  exhibits a strong emission peak around 400–550 nm, which is indicative of the rapid recombination of photogenerated charge carriers. Notably, the PL intensity of the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite is significantly quenched compared to pure  $\text{g-C}_3\text{N}_4$ . This pronounced reduction in emission intensity suggests that the incorporation of  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  effectively suppresses the recombination of charge carriers. As an efficient electron acceptor,  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  facilitates charge transfer and enhances charge separation at the heterojunction interface with  $\text{g-C}_3\text{N}_4$ . In summary, the PL analysis confirms that the formation of the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite improves electron–hole separation and reduces charge recombination, thereby enhancing the photocatalytic performance by making more charge carriers available for surface reactions. This property broadens the potential applications of the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite in the field of photocatalysis.

### 3.3. Photocatalytic reaction mechanism

To investigate the photocatalytic reaction mechanism, various scavengers were employed to identify the main reactive species involved. Tert-butyl alcohol (TBA) was used as a scavenger for hydroxyl radicals ( $\cdot\text{OH}$ ), 1,4-benzoquinone (BQ) for superoxide radical anions ( $\cdot\text{O}_2^-$ ), ammonium oxalate (AO) for photogenerated holes ( $h^+$ ), and dimethyl sulfoxide (DMSO) for photogenerated electrons ( $e^-$ ).<sup>18</sup> A quenching agent solution with an initial

concentration of 10 mmol/L and a volume of 2 mL was added at the onset of light irradiation. As shown in Fig. 5a, the presence of TBA and AO led to a noticeable decrease in the photocatalytic degradation of MB by the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite, compared to the control sample without any scavenger. This can be attributed to the quenching effects of ammonium oxalate (AO) on photogenerated holes ( $\text{h}^+$ ) and of 1,4-benzoquinone (BQ) on superoxide radical anions ( $\cdot\text{O}_2^-$ ), which significantly suppressed the activity of these reactive species in solution and thus



**Figure 5.** (a) Photocatalytic degradation of MB on  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  in the presence of various scavengers, and (b) Proposed Z-scheme mechanism for the photocatalytic activity of the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite

The valence band (VB) and conduction band (CB) positions of  $\text{g-C}_3\text{N}_4$  and  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  can be estimated using the following equations:

$$E_{\text{VB}} = \chi - E^{\circ} + 0.5E_{\text{g}}$$

$$E_{\text{CB}} = E_{\text{VB}} - E_{\text{g}}$$

where  $\chi$  is the absolute electronegativity of the semiconductor (defined as the arithmetic mean of the electron affinity and the first ionization energy of the constituent atoms),  $E^{\circ}$  is the energy of free electrons on the hydrogen scale (4.5 eV), and  $E_{\text{g}}$  is the band gap energy of the semiconductor.

According to Wu et al.,<sup>19</sup> the absolute electronegativity ( $\chi$ ) of  $\text{g-C}_3\text{N}_4$  is 4.72 eV. Based on Mulliken's theory, the electronegativity  $\chi$  of a semiconductor can be accurately calculated using the following formula:

$$\chi = E_{\text{vacuum}} - \frac{E_{\text{g}}}{2}$$

Applied to  $\text{Co}_2[\text{Fe}(\text{CN})_6]$ :

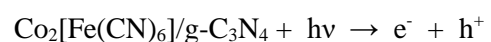
+  $E_{\text{g}} \approx 2.66 \text{ eV}$  (as determined from the UV-Vis DRS spectra);

+  $E_{\text{vacuum}} \approx 5.5 \text{ eV}$ .<sup>20</sup>

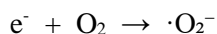
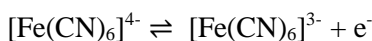
The calculated absolute electronegativity ( $\chi$ ) is 4.17 eV.

Based on the calculations, the conduction band (CB) and valence band (VB) edge potentials of  $\text{g-C}_3\text{N}_4$

are estimated to be  $-1.14 \text{ eV}$  and  $+1.58 \text{ eV}$ , respectively. Similarly, the CB and VB edge potentials of  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  are calculated to be  $-1.66 \text{ eV}$  and  $+1.00 \text{ eV}$ , respectively. On this basis, the separation of photogenerated electron-hole pairs and the formation of reactive radicals in the  $\text{Co}_2[\text{Fe}(\text{CN})_6]/\text{g-C}_3\text{N}_4$  composite under illumination are illustrated in Fig. 5b (Z-scheme). Under light irradiation, both semiconductor materials are photoexcited and generate electron-hole pairs. The photogenerated electrons in the conduction band (CB) of  $\text{g-C}_3\text{N}_4$  tend to recombine with the holes in the valence band (VB) of  $\text{Co}_2[\text{Fe}(\text{CN})_6]$ . This recombination occurs at the contact interface, leading to the accumulation of electrons in the conduction band (CB) of  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  and holes in the valence band (VB) of  $\text{g-C}_3\text{N}_4$ . Consequently, the photogenerated charge carriers are spatially separated across the two semiconductors, while maintaining strong redox potentials. The electrons in the conduction band (CB) of  $\text{Co}_2[\text{Fe}(\text{CN})_6]$  reduce  $\text{O}_2$  to superoxide radicals ( $\cdot\text{O}_2^-$ ), which subsequently contribute to the degradation of MB. In contrast, the holes in the valence band (VB) of  $\text{g-C}_3\text{N}_4$ , due to their insufficient oxidation potential, are unable to oxidize  $\text{H}_2\text{O}$  to generate hydroxyl radicals ( $\cdot\text{OH}$ ), but can directly oxidize MB molecules. Accordingly, the photocatalytic degradation of MB may involve the following processes:







This mechanism satisfactorily explains the reduced recombination of photogenerated electron-hole pairs on the g-C<sub>3</sub>N<sub>4</sub> component of the Co<sub>2</sub>[Fe(CN)<sub>6</sub>]/g-C<sub>3</sub>N<sub>4</sub> composite, as confirmed by the photoluminescence (PL) analysis (Fig. 4d) and the photocatalytic performance results (Fig. 5a). This suppression of charge carrier recombination subsequently enhances the photocatalytic degradation efficiency of MB by the composite under light irradiation.

#### 4. CONCLUSIONS

Nano-sized Co<sub>2</sub>[Fe(CN)<sub>6</sub>] and the Co<sub>2</sub>[Fe(CN)<sub>6</sub>]/g-C<sub>3</sub>N<sub>4</sub> composite were successfully synthesized. The corresponding band gap energies were determined to be 2.66 eV for Co<sub>2</sub>[Fe(CN)<sub>6</sub>], 2.72 eV for g-C<sub>3</sub>N<sub>4</sub>, and 2.71 eV for the composite. Under natural sunlight irradiation for 210 minutes, the Co<sub>2</sub>[Fe(CN)<sub>6</sub>]/g-C<sub>3</sub>N<sub>4</sub> composite exhibited a photocatalytic degradation efficiency of 93.38% for methylene blue in aqueous solution, significantly outperforming the individual components Co<sub>2</sub>[Fe(CN)<sub>6</sub>] and g-C<sub>3</sub>N<sub>4</sub>. The enhanced photocatalytic activity of the composite is mainly attributed to the effective suppression of photogenerated electron-hole recombination. A Z-scheme mechanism for the photocatalytic degradation of methylene blue over the Co<sub>2</sub>[Fe(CN)<sub>6</sub>]/g-C<sub>3</sub>N<sub>4</sub> composite was proposed and discussed in detail by the authors.

#### Acknowledgments

#### REFERENCES

1. Ö. Arar, Ü. Yüksel, N. Kabay, M. Yüksel, Various applications of electrodeionization (EDI) method for water treatment-A short review, *Desalination*, **2014**, 342, 16–22.
2. É. N. Santos, Z. László, C. Hodúr, G. Arthanareeswaran, Photocatalytic membrane filtration and its advantages over conventional approaches in the treatment of oily wastewater: A review, *Asia-Pacific Journal of Chemical Engineering*, **2020**, 2533, 11–9.
3. Lijuan Jiang, Yajun Wang, Changgen Feng, Application of photocatalytic technology in environmental safety, *Procedia Engineering*, **2012**, 45, 993–997.
4. D. Charu, K.P. Sanjay, K. Manmohan, C.T. Subhash and N.B. Parma, Potassium cobalthexacyanoferrate-gel beads for cesium removal: kinetics and sorption studies, *RSC advances*, **2013**, 3, 22102–22110.
5. K. A. Lin, B. Chen, C. Chen, Evaluating Prussian Blue Analogues M II 3[MIII(CN)<sub>6</sub>]<sub>2</sub> (MII = Co, Cu, Fe, Mn, Ni; MIII = Co, Fe) as Activators for Peroxymonosulfate in water, *RSC Advances*, **2016**, 95, 1–36.
6. F. Karadas, H. El-Faki, E. Deniz, C.T. Yavuz, S. Aparicio, M. Atilhan, CO<sub>2</sub> adsorption studies on Prussian blue analogues, *Microporous and Mesoporous Materials*, **2012**, 162, 91–97.
7. X. Li, J. Liu, A.I. Rykov, H. Han, C. Jin, X. Liu, J. Wang, Excellent photo-Fenton catalysts of Fe-Co Prussian blue analogues and their reaction mechanism study, *Applied Catalysis B: Environmental*, **2015**, 179, 196–205.
8. Lan H.T.L., Son A.N., Trung D.N., Van C.T.L, Hai V.C, Ngoc B.N. and Thao P.T.L, Prussian Blue Analogues of A<sub>2</sub>[Fe(CN)<sub>6</sub>] (A: Cu<sup>2+</sup>, Co<sup>2+</sup>, and Ni<sup>2+</sup>) and Their Composition-Dependent Sorption Performances towards Cs<sup>+</sup>, Sr<sup>2+</sup>, and Co<sup>2+</sup>, *Journal of Nanomaterials*, **2021**, 20213, 1–12.
9. S. Adak, L.L. Daemen, M. Hartl, D. Williams, J. Summerhill, H. Nakotte, Thermal expansion in 3d-metal Prussian Blue Analogs—A survey study, *Journal of Solid State Chemistry*, **2011**, 184, 2854–2861.
10. S. Xu, X. Qian, G. Li, Size and morphology-controlled Ni<sub>2</sub>[Fe(CN)<sub>6</sub>]<sub>x</sub>H<sub>2</sub>O Prussian Blue analogue fabricated via a hydrothermal route, *Materials Research Bulletin*, **2008**, 43, 135–140.
11. E. Jang, D. W. Kim, S. H. Hong, Y. M. Park, T. J. Park, Visible light-driven g-C<sub>3</sub>N<sub>4</sub>@ZnO heterojunction photocatalyst synthesized via atomic layer deposition with a specially designed rotary reactor, *Applied Surface Science*, **2019**, 487, 206–210.
12. M. Kim, S. Hwang and J.S. Yu, Novel ordered nanoporous graphitic carbon nitride with g-C<sub>3</sub>N<sub>4</sub> stoichiometry as a support for Pt-Ru anode catalyst in DMFC, *J. Mater. Chem.*, **2007**, 17, 1656–1659.
13. N. V. Kim, N. T. V. Nga, T. T. T. Phuong, N. L. Tuan, V. Vien, Synthesis of g-C<sub>3</sub>N<sub>4</sub>/ZnO composites with enhanced photocatalytic activity under visible light, *Vietnam Journal of Chemistry*, **2018**, 56, 220–225.
14. F. Dong, Y. Li, Z. Wang, W.-K.Ho, Enhanced visible light photocatalytic activity and oxidation ability of porous graphene-like g-C<sub>3</sub>N<sub>4</sub> nanosheets via thermal exfoliation, *Applied Surface Science*, **2015**, 358, 393–403.
15. X. Li, J. Liu, A.I. Rykov, H. Han, C. Jin, X. Liu, & J. Wang, Excellent photo-Fenton catalysts of Fe-Co Prussian blue analogues and their reaction



mechanism study, *Applied Catalysis B: Environmental*, **2015**, 179, 196–205.

16. Y. Chen, R.C. Zhou, Y.H. Wan, J.W. Hao, H.R. You, X.M. Liu, H. Yang,  $\text{KCo}_x\text{Mn}_{1-x}[\text{Fe}(\text{CN})_6]$ /Carbon nanotube composite as high capacity anode for Li-ion batteries, *Journal of Electroanalytical Chemistry*, **2021**, 887, 115151.
17. S. Jiao, J. Tuo, H. Xie, Z. Cai, S. Wang, J. Zhu, The electrochemical performance of  $\text{Cu}_3[\text{Fe}(\text{CN})_6]_2$  as a cathode material for sodium-ion batteries, *Materials Research Bulletin*, **2017**, 86, 194–200.
18. L. Shi, L. Liang, J. Ma, F. Wang, & J. Sun, Remarkably enhanced photocatalytic activity of ordered mesoporous carbon/g- $\text{C}_3\text{N}_4$  composite photocatalysts under visible light, *Dalton Transactions*, **2014**, 43, 7236–7244.
19. Y. Wu, L. Tao, J. Zhao, X. Yue, Y. Li, Chuanyi Wang,  $\text{TiO}_2/\text{g-C}_3\text{N}_4$  nanosheets hybrid photocatalyst with enhanced photocatalytic activity under visible light irradiation, *Research on Chemical Intermediates*, **2016**, 42, 3609–3624.
20. L.C. Wang, P.Y. Chiou, Y.P. Hsu, C.L. Lee, C.H. Hung, Y.H. Wu, W.J. Wang, G.L. Hsieh, Y.C. Chen, L.C. Chang, W.P. Su, D. Manoharan, M.C. Liao, S. Thangudu, W.P. Li, C.H. Su, H.K. Tian & C.S. Yeh, *Nature Communications*, **2023**, 14, 4709.